

An accurate theoretical description of fluids composed of fully anisotropic molecules: Application to C_{2v} symmetry

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We use two molecular integral equation approximations to compute the thermodynamic properties and microscopic structure of two liquids composed of planar molecules with C_{2v} symmetry, namely SO_2 and H_2S . These approximations couple the exact molecular Ornstein–Zernike equation with the hypernetted chain (HNC) and reference-hypernetted chain (RHNC) closures. The theoretical results obtained for various thermodynamic states agree remarkably well with molecular dynamics calculations. In particular, the atom-atom distribution functions are very well reproduced. We find that the RHNC approximation with a hard-sphere fluid reference system offers notable improvement over HNC in the pressure calculation. We include also a self-consistent mean field calculation to incorporate the effect of polarizability on the dielectric constant of liquid SO_2 . Final results for this quantity are in excellent agreement with experimental values. In contrast, the model used for the electrostatic interactions in H_2S leads to anomalously high permanent dipole moments, compared to experiment, and consequently to dielectric constants that are completely off the experimental data.

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I. INTRODUCTION

In a recent paper, Lado *et al.*¹ presented an algorithm for the numerical solution of Ornstein–Zernike (OZ) integral equations for fluids composed of fully anisotropic particles. As an illustration of the method, they solved the hypernetted chain (HNC) equation for a simple model of liquid SO_2 proposed by Sokolić *et al.*,² whereby sulphur dioxide molecules are represented by a three-center Lennard–Jones potential with C_{2v} symmetry without electrostatic interactions.

The results of Ref. 1, although taken at relatively high temperature and moderate density, clearly suggested that the proposed algorithm and the HNC approximation are able to produce a reasonable description of not-so-simple fluids. It was also clear, however, that further studies were needed to obtain a better assessment of this approach. This is consequently the main goal of the present work. On the one hand, we seek here to examine thermodynamic states that are more representative of the liquid phase (i.e., lower temperatures and higher densities) and more realistic potential models that include electrostatic interactions. For this, we have focused again on the model used in Ref. 1 together with a point charge model also proposed by Sokolić *et al.*² On the other hand, we also consider a somewhat less anisotropic system proposed by Forester *et al.*³ to model a different molecular liquid, namely H_2S . In this case, we use a three-site point charge potential with a single Lennard–Jones center. For these models, we have performed calculations on the liquid side of the phase diagram in the vicinity of the zero-pressure line, which we know to be a good low-temperature estimate

of the liquid branch of the gas–liquid equilibrium curve. It is shown in this work that, good as they seem, the HNC results leave room for improvement, in particular with regard to the pressure, which is typically overestimated. We have thus carried out reference-hypernetted chain (RHNC)⁴ calculations with a hard-sphere reference fluid whose sphere diameter is optimized according to the free energy minimization criterion.⁵ A significant improvement is thereby achieved, especially for the less anisotropic of the two substances, H_2S .

Comparison with simulation results of the spherical harmonic projections of the molecular pair distribution function calculated in Ref. 1 showed a surprising ability of the HNC approximation to describe the averaged microscopic ordering of the fully anisotropic molecules. However, this type of representation, beyond a few low-order coefficients, has a difficult physical interpretation. A more intuitive picture of the microstructure is often provided by atom-atom distribution functions. Thus another task we undertake here is a straightforward extension of the algorithm proposed by Alvarez *et al.*,⁶ in the context of fluids composed of linear molecules, to calculate the atom-atom pair distribution functions by multidimensional integration over the full molecular pair distribution function. We find that, except for some fine details, both the HNC and the RHNC approximations accurately reproduce the simulated atom-atom correlations.

Finally, an important quantity that can be evaluated from the microscopic information provided by the integral equation calculations is the dielectric constant. Comparison with dielectric constants from simulation, however, is plagued with the numerous uncertainties associated with the determination of this quantity by simulation, which typically requires extremely long runs⁷ and is very sensitive to system

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size. Encouraged by previous experience for different substances like HCl (Ref. 8) and CH₃CN (Ref. 9), we decided instead to compare our results with purely experimental data. In order to do this, the effect of polarizability has to be taken into account. This is done using the self-consistent mean field approach (SCMF) proposed by Carnie and Patey.¹⁰ We find that the results obtained in this way for liquid SO₂, both in HNC and RHNC approximation, agree well with the experimental dielectric constants. In contrast, we conclude that the manner in which the electrostatic interactions are modeled in H₂S,³ while leading to reasonable thermodynamic properties, produces an anomalously high dipole moment which in turn is responsible for excessively high values of the dielectric constant obtained, even without inclusion of polarizability effects. Aside from these modeling problems, the theory performs remarkably well for all the thermodynamic properties and states considered.

The remainder of the paper proceeds along the following lines. The next section offers a brief summary of the elements of the theory required for the calculations carried out in this work. The models and thermodynamic states on which we have focused are presented in Sec. III. In Sec. IV we summarize the main results and conclusions.

II. A BRIEF SKETCH OF THE THEORY

An essential part of the calculations presented in this paper consists of the OZ equation for molecular fluids,

$$\gamma(12) = \frac{\rho}{8\pi^2} \int d\mathbf{3}c(13)[c(32) + \gamma(32)], \quad (1)$$

where ρ is the number density and $\int d\mathbf{3} = \int d\mathbf{r}_3 d\omega_3$ denotes integration over particle positions and orientations. This equation forms a closed set together with a general closure relation resulting from a diagrammatic analysis which relates the structure of the liquid and the pair potential $u(12)$,

$$c(12) = \exp[-\beta u(12) + \gamma(12) + B(12)] - \gamma(12) - 1, \quad (2)$$

where $c(12)$ is the direct correlation function and $\gamma(12) = g(12) - 1 - c(12)$ is the indirect correlation function, with $g(12)$ the molecular pair distribution function. $B(12)$ is the so-called bridge function, which formally sums a series of complex terms in a diagrammatic density expansion. Here we will consider two approximations to this quantity: $B(12) = 0$, the hypernetted chain approximation, and $B(12) = B_{\text{HS}}(r; d)$, the reference-hypernetted chain approximation,⁴ where B_{HS} is the bridge function of a reference hard-sphere fluid, calculated with the Verlet–Weis–Grudke–Henderson^{11,12} parametrization. The hard-sphere diameter d is chosen so as to minimize the free energy,⁵ which results in the condition

$$\rho \int d\mathbf{r} [g_{000}^{\text{HS}}(r) - g_{\text{HS}}(r; d)] \frac{\partial B_{\text{HS}}(r; d)}{\partial d} = 0, \quad (3)$$

where $g_{000}^{\text{HS}}(r)$ is the center-to-center distribution function of the fluid and $g_{\text{HS}}(r; d)$ that of the reference hard-sphere system.

As noted in Ref. 1, the key to dealing with Eqs. (1) and (2) is the expansion of all angular functions in generalized spherical harmonics, i.e.,

$$X(12) = \sum_{l_1, l_2, m, n_1, n_2} [(2l_1 + 1)(2l_2 + 1)]^{1/2} \times X_{l_1 l_2 m}^{n_1 n_2}(r_{12}) D_{m n_1}^{l_1}(\omega_1) * D_{m n_2}^{l_2}(\omega_2)^*, \quad (4)$$

where, as in Ref. 1, $\omega = (\phi, \theta, \chi)$ are the Euler angles referred to the axial line \mathbf{r}_{12} between molecular centers and $D_{mn}^l(\omega)$ is a generalized spherical harmonic in the convention of Rose.¹³ One sees immediately that $g_{000}^{\text{HS}}(r)$ in Eq. (3) is the first term in the expansion of $g(12)$.

Details of the algorithm used to solve Eqs. (1) and (2) in terms of the expansion coefficients of the functions involved can be found in Ref. 1. Once the set of nonlinear equations is solved we can proceed to calculate both thermodynamic and dielectric properties. Thus we have

$$\frac{\beta U}{N} = 2\pi\rho \int_0^\infty \langle \beta u(12)g(12) \rangle_{\omega_1 \omega_2} r_{12}^2 dr_{12} \quad (5)$$

and

$$\frac{\beta P}{\rho} = 1 - \frac{2}{3} \pi\rho \int_0^\infty \left\langle r_{12} \frac{\partial \beta u(12)}{\partial r_{12}} g(12) \right\rangle_{\omega_1 \omega_2} r_{12}^2 dr_{12}, \quad (6)$$

where

$$g(12) = \exp[-\beta u(12) + \gamma(12) + B(12)] \quad (7)$$

and the angular brackets denote averages over all angles. In the expressions above, we have put as usual $\beta = 1/k_B T$, where T is the absolute temperature and k_B is Boltzmann's constant. Note that $g(12)$ is constructed indirectly from the summation of $\gamma(12)$ coefficients, rather than directly from those of $g(12)$ itself; this avoids the well-known poor convergence of the spherical harmonic series of $g(12)$ for anisotropic molecules. The reconstruction of $\gamma(12)$ from expansion (4) is known to be rapidly convergent and well behaved.

The atom-atom correlation functions are similarly obtained by integration over $g(12)$; in numerical format, we write

$$g_{\alpha\beta}(r) = w^3 \sum_{i=1}^{n_r} \sum_{i_1, i_2, j, k_1, k_2=1}^n w_{i_1} w_{i_2} g(12) \left(\frac{R_i}{r} \right)^2 \times \Delta R \delta[r - \mathfrak{R}_{\alpha\beta}(R_i, x_{i_1}, x_{i_2}, y_j, z_{k_1}, z_{k_2})], \quad (8)$$

where $\mathfrak{R}_{\alpha\beta}(R_i, x_{i_1}, x_{i_2}, y_j, z_{k_1}, z_{k_2})$ is the function that gives the atom-atom distance between sites α and β for a given center-to-center separation R_i and relative orientations described by $x_{i_m} = \cos \theta_{i_m}$, $y_j = \cos \phi_j$, and $z_{k_m} = \cos \chi_{k_m}$, where $\theta_1, \theta_2, \chi_1, \chi_2$, $\phi = \phi_1 - \phi_2$ is the set of Euler angles required to describe the relative orientation of two nonlinear molecules. The w_{i_m} in Eq. (8) are the weights of the Gauss–Legendre quadrature for the points x_{i_m} and $w = 1/n$ is the constant weight of Gauss–Chebyshev quadrature for the points y_j and z_{k_m} , with n the number of grid points used in

each angular integration and n_r , the number of grid points on the center-to-center R axis, tabulated in equal size intervals ΔR .

The dielectric constant can be obtained from the Kirkwood formula,

$$\frac{(\epsilon-1)(2\epsilon+1)}{9\epsilon} = y g_K, \quad (9)$$

where $y = 4\pi\beta\rho\mu^2/9$, μ is the dipole moment, and g_K the Kirkwood g factor, a well-known order parameter that can be expressed in terms of correlation functions as

$$g_K = 1 + \frac{4}{3}\pi\rho \int_0^\infty r^2 g(r;00,110) dr. \quad (10)$$

Here we have used the expansion coefficient $g(r;00,110)$ of $g(12)$ in the space-fixed reference frame.¹⁴

To incorporate the effects of polarizability by means of the SCMF approach,¹⁰ we recall that the effective dipole moment μ_{eff} is related to the permanent one μ by

$$\beta\mu_{\text{eff}}^2 = \frac{\beta\mu^2}{(1+\alpha c)^2} + \frac{3\alpha}{1+\alpha c}, \quad (11)$$

where α is the spherical polarizability¹⁵ and

$$c = c(\mu_{\text{eff}}) = \frac{2U_{\text{DD}}(\mu_{\text{eff}}) + U_{\text{DQ}}(\mu_{\text{eff}})}{N\mu_{\text{eff}}^2}. \quad (12)$$

Here U_{DD} and U_{DQ} are, respectively, the dipole-dipole and dipole-quadrupole contribution to the internal energy for a fluid with a dipole moment μ_{eff} . These quantities can be easily evaluated once $g(12)$ is known by replacing $u(12)$ in Eq. (5) with its dipole-dipole and dipole-quadrupole terms. The standard procedure involves carrying out integral equation calculations for various values of μ_{eff} in the vicinity of the model dipole moment μ_0 , consistent with the initial charge distribution. One then plots μ_{eff}^2 vs μ^2 , the latter obtained from Eqs. (11) and (12) (typically a linear relation), and from this graph interpolates the value of μ_{eff}^2 that is consistent with the desired permanent dipole moment of the nonpolarizable system, μ_0 . The calculations are then repeated for this effective dipole moment and the dielectric constant is calculated from

$$\frac{(\epsilon-1)(2\epsilon+1)}{9\epsilon} - \frac{(\epsilon_\infty-1)(2\epsilon_\infty+1)}{9\epsilon_\infty} = y_{\text{eff}} g_K(\mu_{\text{eff}}), \quad (13)$$

where $y_{\text{eff}} = 4\pi\beta\mu_{\text{eff}}^2\rho/9$ and ϵ_∞ is the high-frequency dielectric constant, for which we use the Clausius–Mossotti value.

III. RESULTS AND DISCUSSION

The geometrical and potential energy parameters for the two models used to represent SO_2 and for the point-charge model of H_2S are collected in Fig. 1 and Table I. Cross interactions are evaluated by means of the Lorentz–Berthelot mixing rules; charges are given in terms of the elementary electronic charge e . For SO_2 , we have studied a thermodynamic state corresponding to $T = 249.5$ K and $V = 42.94$

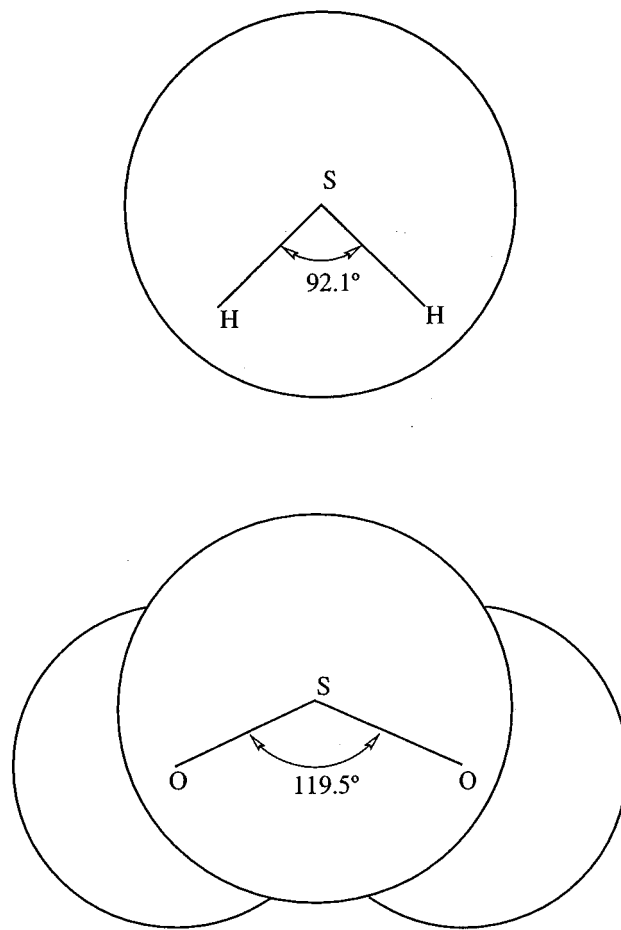


FIG. 1. Molecular geometries of H_2S (top) and SO_2 (bottom) used in this work.

$\text{cm}^3 \text{mole}^{-1}$ ($\rho = 0.0140 \text{ \AA}^{-3}$), and for H_2S , $T = 293$ K and $V = 43.10 \text{ cm}^3 \text{mole}^{-1}$ ($\rho = 0.0139 \text{ \AA}^{-3}$). We note that the charge distribution of model B for SO_2 leads to a dipole moment $\mu_0 = 1.640$ D, close to the experimental gas phase value $\mu_{\text{exp}} = 1.634$ D, whereas the charge distribution in our model for H_2S leads to $\mu_0 = 2.087$ D, which is more than double the experimental gas phase value of $\mu_{\text{exp}} = 0.974$ D. Forester *et al.*³ have proposed a more elaborate model that incorporates additional auxiliary charges and reduces considerably the net dipole moment while still producing reasonable results for the thermodynamic quantities. Nevertheless, to limit the complexity of the model while retaining the correct molecular shape, we will continue to use the model described in Table I. We will see later, however, that the artificially high dipole moment yields essentially meaningless dielectric properties.

With these considerations in mind, we have numerically solved the HNC and the optimized RHNC equations. In Fig. 2 we show the atom-atom pair distribution for model A of SO_2 . One sees that both the HNC and RHNC approximations yield good agreement with molecular dynamics results.¹⁶ Only in the g_{ss} correlation is the HNC equation slightly out of phase, which is corrected by the RHNC equation. The

TABLE I. Potential parameters for the molecular liquid models considered in this work.

SO ₂		
Model A	Model B	H ₂ S
$\epsilon_{S-S}=182$ K; $\epsilon_{O-O}=72$ K	$\epsilon_{S-S}=154.5$ K; $\epsilon_{O-O}=62.3$ K	$\epsilon_{S-S}=125$ K
$\sigma_{S-S}=3.61$ Å; $\sigma_{O-O}=3$ Å	$\sigma_{S-S}=3.585$ Å; $\sigma_{O-O}=2.993$ Å	$\sigma_{S-S}=3.71$ Å
	$q_S=+0.470e$; $q_O=-0.235e$	$q_S=+0.470e$; $q_H=-0.235e$

latter approximation, however, tends to overestimate the first peak, probably due to the limitations of the spherically symmetric reference system. The same comments can be made about the results for model B, shown in Fig. 3. Thermodynamic results are given in Table II, specifically the internal energy and pressure for these models from both the HNC and RHNC approximations and from molecular dynamics (MD) simulations. We find that the internal energies are in

good agreement with MD for both approximations. For the compressibility factor, comparison with simulation shows that, as usual, the HNC equation overestimates this quantity, while here the improvement introduced by the spherically symmetric reference system falls short of what is needed. We have attempted to improve upon this by incorporating a non-spherical reference system with a bridge function evaluated using Verlet's modified approximation, along the lines explored in earlier works dealing with molecular fluids.¹⁷ Unfortunately, the appearance of a pole in the denominator of

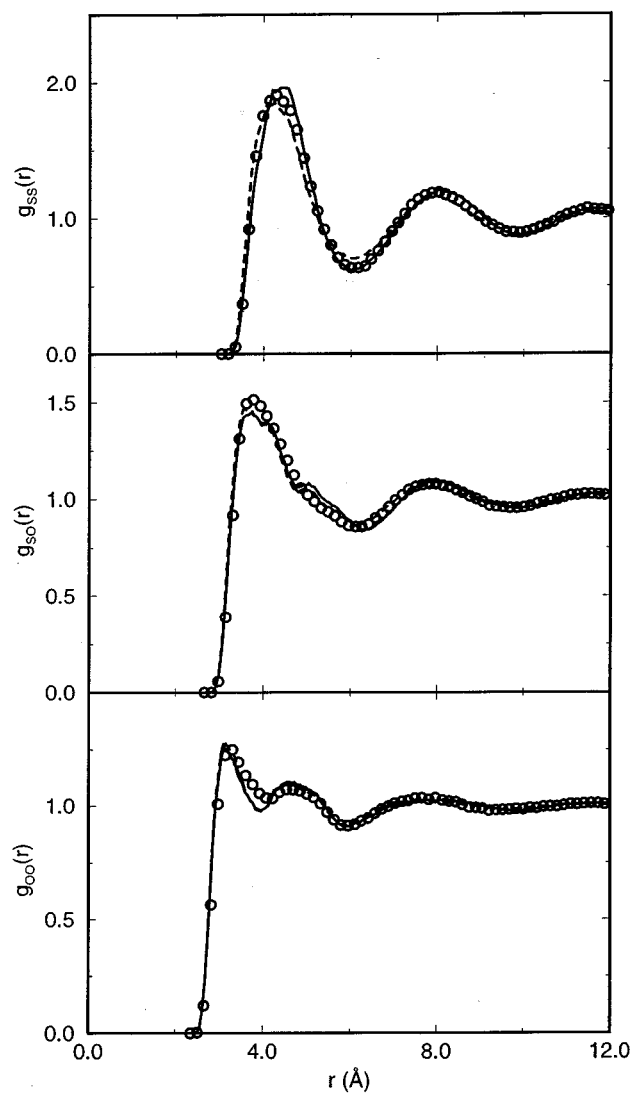


FIG. 2. Atom-atom distribution functions $g_{\alpha\beta}(r)$ for model A of liquid SO₂ at $T=249.5$ K. Computed RHNC (continuous line) and HNC (dashed line) results vs MD simulation data (open circles). See Table I for model parameters.

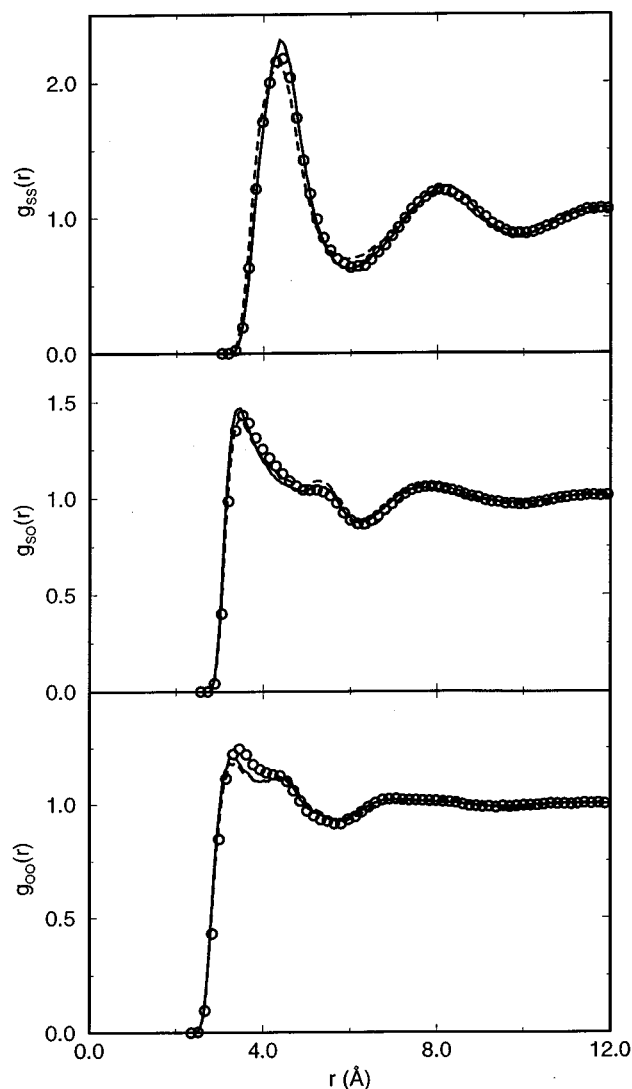


FIG. 3. Same as Fig. 2 for model B of liquid SO₂.

TABLE II. Internal energy and compressibility factor from theory and MD simulation, together with the experimental values taken from Refs. 2 and 3.

	SO ₂		H ₂ S
	Model A	Model B	
$(\beta U/N)_{\text{exp.}}$		-11.51	-5.32
$(\beta U/N)_{\text{MD}}$	-11.55 ± 0.03	-11.34 ± 0.02	-5.24 ± 0.06
$(\beta U/N)_{\text{HNC}}$	-11.10	-11.20	-5.12
$(\beta U/N)_{\text{RHNC}}$	-11.26	-11.37	-5.15
$(\beta P/\rho)_{\text{exp.}}$		0.001	0.18
$(\beta P/\rho)_{\text{MD}}$	0.1 ± 0.2	0.2 ± 0.4	0.4 ± 0.4
$(\beta P/\rho)_{\text{HNC}}$	4.20	3.81	1.05
$(\beta P/\rho)_{\text{RHNC}}$	2.60	1.90	0.55

the Verlet bridge function prevents the calculation of this reference system for the cases of interest in this work. Thus improving upon the hard-sphere bridge function for the type of systems considered here remains an open question.

For liquid H₂S, the smaller anisotropy of the molecule (see Fig. 1) improves the situation. The site-site distribution functions, depicted in Fig. 4, obtained from the HNC and RHNC equations are essentially indistinguishable and both are in excellent agreement with simulation. Table II shows that agreement of the calculated compressibility factor is better than that for SO₂, especially for the RHNC result, while the internal energy values remain good.

If one is now willing to compare the computed dielectric constants with *experimental* values, the situation is quite different for the two substances. For liquid SO₂ at the thermodynamic state specified earlier, the experimental value is $\epsilon_{\text{exp}} = 18.7$, while the theoretical estimates (without inclusion of polarizability) are $\epsilon_{\text{HNC}} = 10.0$ and $\epsilon_{\text{RHNC}} = 11.0$. An estimate from a 40 000 step MD run gives $\epsilon_{\text{MD}} = 9.5 \pm 1.0$, in good agreement with these results. The differences between experimental and model results suggest that polarizability should play an important role here, which can be readily accepted by noting that the experimental value of the spherical polarizability is $\alpha \approx 4.28 \text{ \AA}^{-3}$. In contrast, the theoretical calculations for liquid H₂S yield $\epsilon_{\text{HNC}} = 16.8$ and $\epsilon_{\text{RHNC}} = 17.1$, which compare poorly with the experimental value of $\epsilon_{\text{exp}} \approx 6$. As mentioned earlier, this can be understood as a consequence of the anomalously high dipole moment (more than twice the experimental value) derived from our model charge distribution. Trying to incorporate polarizability into the present calculations for H₂S is not sensible.

For liquid SO₂ we perform calculations at various effective dipole moments and thus, by means of Eqs. (11) and (12), obtain the representation of μ_{eff}^2 vs μ^2 shown in Fig. 5. The well-known linear form of this dependence makes it easy to extract the value of μ_{eff} that corresponds to $\mu_0 = 1.64$ D ($\mu^* = 1.693$); this turns out to be $\mu_{\text{eff}} = 1.88$ D ($\mu^* = 1.946$). These calculations have been performed in HNC approximation; the RHNC equation leads to very similar values of the effective dipole moments. With this μ_{eff} , the new values of the dielectric constant obtained from Eq. (13) are $\epsilon_{\text{HNC}} = 18.0$ and $\epsilon_{\text{RHNC}} = 20.1$, in good agreement with the experimental $\epsilon_{\text{exp}} = 18.7$, as was the case earlier for HCl (Ref. 8) and CH₃CN (Ref. 9). The SCMF calculation

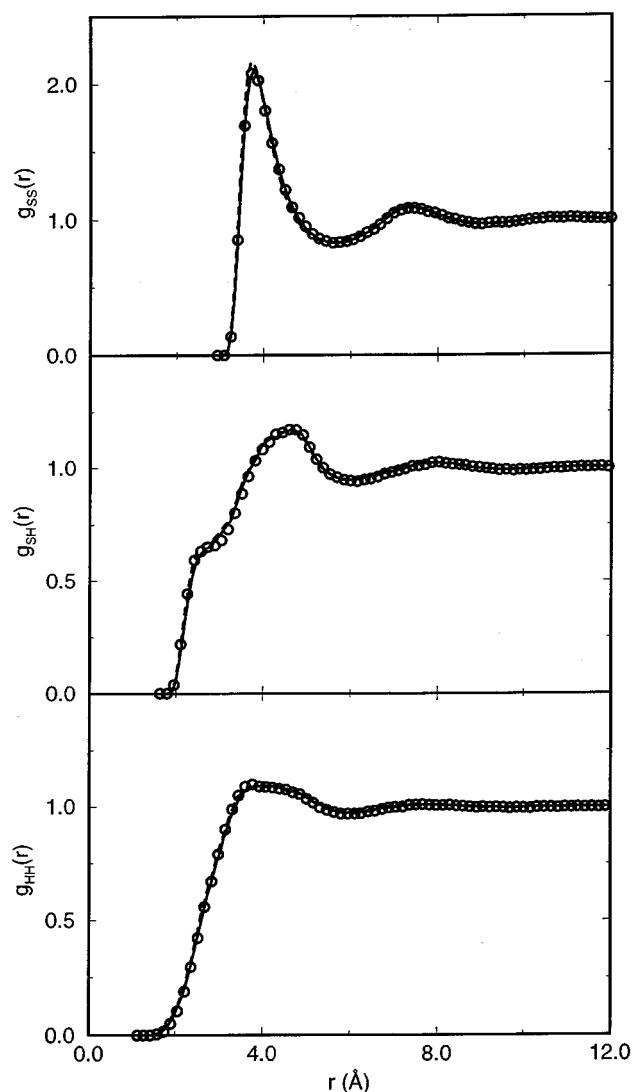


FIG. 4. Atom-atom distribution functions $g_{\alpha\beta}(r)$ for the model of liquid H₂S at $T = 293$ K used in this work. Computed RHNC (continuous line) and HNC (dashed line) results vs MD simulation data (open circles). Results for $g_{\alpha\beta}(r)$ from the two theoretical approximations are practically indistinguishable on the scale of the figure. See Table I for model parameters.

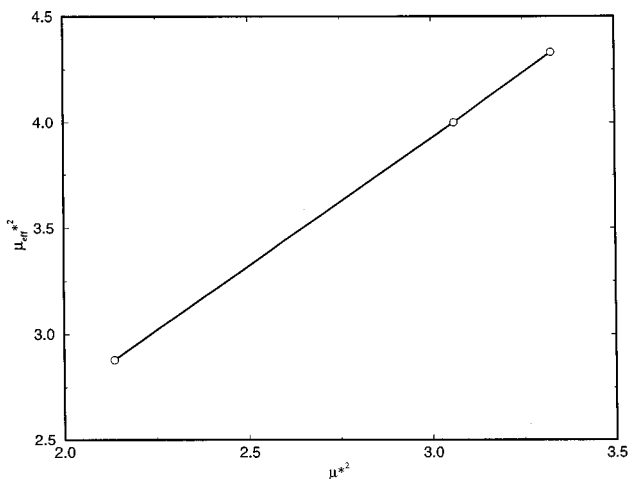


FIG. 5. Calculated values of μ_{eff}^{*2} vs μ^{*2} for model B of liquid SO₂. The continuous line is merely a guide to the eye.

combined with HNC or RHNC results is thus found to be a useful tool with remarkable predictive power so far as the dielectric constant is concerned.

IV. SUMMARY AND CONCLUSIONS

In summary, we have presented another application of a powerful molecular integral equation implementation for the description of fluids composed of fully anisotropic particles with a C_{2v} symmetry that is shared by numerous substances of interest. Computed results for both structure and thermodynamics are in good accord with molecular dynamics calculations. The use of a spherical reference system with the RHNC equation is particularly good for H_2S , which is quite encouraging, given the similarities between the model used here to describe this substance and those used for liquid water consistent with the C_{2v} symmetry, both with three (SPC) and four sites (TIP4P). Later work will be devoted to an extensive integral equation study of these rigid models of water.

The SCMF calculations carried out for liquid SO_2 lead to dielectric constants that closely reproduce the experimental values. The model used for H_2S , however, proves to be unsuitable for the calculation of dielectric properties, given its inadequate representation of the electrostatic components of the interactions.

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